

## Chemical Formula Writing Worksheet

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl<sub>2</sub>", as shown.

	zinc	iron (II)	iron (III)	gallium	silver	lead (IV)
chloride	ZnCl <sub>2</sub>					
acetate						
nitrate						
oxide						
nitride						
sulfate						

Write the formulas for the following compounds:

- 1) copper (II) chloride \_\_\_\_\_
- 2) lithium acetate \_\_\_\_\_
- 3) vanadium (III) selenide \_\_\_\_\_
- 4) manganese (IV) nitride \_\_\_\_\_
- 5) beryllium oxide \_\_\_\_\_
- 6) sodium sulfate \_\_\_\_\_
- 7) aluminum arsenide \_\_\_\_\_
- 8) potassium permanganate \_\_\_\_\_
- 9) chromium (VI) cyanide \_\_\_\_\_
- 10) tin (II) sulfite \_\_\_\_\_
- 11) vanadium (V) fluoride \_\_\_\_\_
- 12) ammonium nitrate \_\_\_\_\_

## Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write "ZnCl<sub>2</sub>", as shown.

	<i>zinc</i>	<i>iron (II)</i>	<i>iron (III)</i>	<i>gallium</i>	<i>silver</i>	<i>lead (IV)</i>
<i>chloride</i>	ZnCl <sub>2</sub>	FeCl <sub>2</sub>	FeCl <sub>3</sub>	GaCl <sub>3</sub>	AgCl	PbCl <sub>4</sub>
<i>acetate</i>	Zn(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>2</sub>	Fe(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>2</sub>	Fe(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>3</sub>	Ga(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>3</sub>	Ag <sub>2</sub> C <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	Pb(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>4</sub>
<i>nitrate</i>	Zn(NO <sub>3</sub> ) <sub>2</sub>	Fe(NO <sub>3</sub> ) <sub>2</sub>	Fe(NO <sub>3</sub> ) <sub>3</sub>	Ga(NO <sub>3</sub> ) <sub>3</sub>	AgNO <sub>3</sub>	Pb(NO <sub>3</sub> ) <sub>4</sub>
<i>oxide</i>	ZnO	FeO	Fe <sub>2</sub> O <sub>3</sub>	Ga <sub>2</sub> O <sub>3</sub>	Ag <sub>2</sub> O	PbO <sub>2</sub>
<i>nitride</i>	Zn <sub>3</sub> N <sub>2</sub>	Fe <sub>3</sub> N <sub>2</sub>	FeN	GaN	Ag <sub>3</sub> N	Pb <sub>3</sub> N <sub>4</sub>
<i>sulfate</i>	ZnSO <sub>4</sub>	FeSO <sub>4</sub>	Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	Ga <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	Ag <sub>2</sub> SO <sub>4</sub>	Pb(SO <sub>4</sub> ) <sub>2</sub>

Write the formulas for the following compounds:

- 1) copper (II) chloride CuCl<sub>2</sub>
- 2) lithium acetate LiC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- 3) vanadium (III) selenide VSe
- 4) manganese (IV) nitride Mn<sub>3</sub>N<sub>4</sub>
- 5) beryllium oxide BeO
- 6) sodium sulfate Na<sub>2</sub>SO<sub>4</sub>
- 7) aluminum arsenide AlAs
- 8) potassium permanganate KMnO<sub>4</sub>
- 9) chromium (VI) cyanide Cr(CN)<sub>6</sub>
- 10) tin (II) sulfite SnSO<sub>3</sub>
- 11) vanadium (V) fluoride VF<sub>5</sub>
- 12) ammonium nitrate NH<sub>4</sub>NO<sub>3</sub>